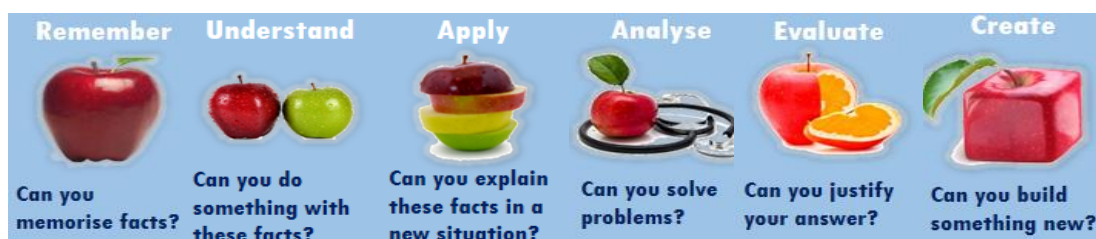




Whickham School - Year 7 Science Homework

C3 Chemical Reactions

	Developing	Securing	Mastering
Chemical Reactions lesson 1	<u>Developing 1</u> 1. Describe what happens when ice melts 2. Describe what happens when you burn magnesium in air.	<u>Securing 1</u> 1. Explain what happens when ice melts 2. Explain what happens when you burn magnesium in air.	<u>Mastering 1</u> 1. Compare and contrast chemical changes with physical changes. Include examples to illustrate your answer (6 marks)
Chemical Reactions lesson 2	<u>Developing 2</u> Write your own sentences to describe each of the following key words. 1. Product 2. Reactant 3. Arrow 4. Word equation 5. Formula equation	<u>Securing 2</u> Put the following in to a paragraph to describe chemical equations 1. Product 2. Reactant 3. Arrow 4. Word equation 5. Formula equation	<u>Mastering 2</u> Demonstrate the meaning of the following terms using an example 1. Product 2. Reactant 3. Arrow 4. Word equation 5. Formula equation
Chemical Reactions lesson 3	<u>Developing 3</u> Produce a leaflet on 'alternative fuels for cars'.	<u>Securing 3</u> Produce a leaflet to compare alternative fuels for cars.	<u>Mastering 3</u> Produce a leaflet on alternative fuels including the social, economic and moral impact of using these fuels
Chemical Reactions lesson 4	<u>Developing 4</u> Carry out research on the compound hydrogen peroxide. 1. How is it used? 2. How is it stored to slow down decomposition reactions?	<u>Securing 4</u> Hydrogen peroxide is a compound. 1. Why is it classed as a compound? 2. What properties does it have?	<u>Mastering 4</u> Hydrogen peroxide has the chemical formulae is H_2O_2 which is very similar to water (H_2O) why do they not have the same properties?
Chemical Reactions lesson 5	<u>Developing 5</u> Calculate the missing mass for these equations $C(s) + O_2(g) \rightarrow CO_2(g)$ 1.2g 3.2g $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$ 5.0g 2.2g	<u>Securing 5</u> Calculate the missing mass for these equations $C(s) + O_2(g) \rightarrow CO_2(g)$ 14mg 3.0 g $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$ 0.05kg 20.2g	<u>Mastering 5</u> Calculate the missing mass for these equations, but you will also need to balance them $CH_4(g) + O_2(g) \rightarrow CO_2(g) + H_2O(g)$ 6.4g 4.4g 3.6g $Fe(s) + O_2(g) \rightarrow Fe_2O_3(s)$ 8.93g 12.77g
Chemical Reactions lesson 6	<u>Developing 6</u> Find pictures to show the different types of reaction	<u>Securing 6</u> Define the different types of reaction Combustion	<u>Mastering 6</u> Compare the different types of reaction Combustion





	<p>introduced in the 'REACTIONS' module. Label each picture.</p> <p>Combustion Thermal decomposition Oxidation Endothermic Decomposition Exothermic Physical change Chemical change Dissolving</p>	<p>Thermal decomposition Oxidation Endothermic Decomposition Exothermic Physical change Chemical change Dissolving</p>	<p>Thermal decomposition Oxidation Endothermic Decomposition Exothermic Physical change Chemical change Dissolving</p>
<p><u>Forward Thinking & FLIP Learning</u></p> <ol style="list-style-type: none"> Research the chemical formula of acids what do they have in common? Research the chemical formula of alkali what do they have in common? 	<p><u>Push yourself - 10 point RED HOT challenge task</u></p> <div style="display: flex; align-items: flex-start;">  <div> <p><u>Popping candyyum yum</u></p> <p>Is this a chemical reaction? How do you know? Research the chemical formulae what elements does it contain? What does it react with in your mouth? Can you write a word equation for this? Is it an exothermic or endothermic reaction?</p> </div> </div>		

